

Europäisches Patentamt
European Patent Office
Office européen des brevets



(11) **EP 1 254 709 A2**

(12) **EUROPEAN PATENT APPLICATION**

(43) Date of publication:
06.11.2002 Bulletin 2002/45

(51) Int Cl.7: **B01J 23/00, B01J 23/20,
B01J 23/28, C07C 253/24,
C07C 51/215**

(21) Application number: **02252866.5**

(22) Date of filing: **23.04.2002**

(84) Designated Contracting States:
**AT BE CH CY DE DK ES FI FR GB GR IE IT LI LU
MC NL PT SE TR**
Designated Extension States:
AL LT LV MK RO SI

(30) Priority: **25.04.2001 US 286217 P**

(71) Applicant: **ROHM AND HAAS COMPANY**
Philadelphia, Pennsylvania 19106-2399 (US)

(72) Inventors:
• **Bogan Jr., Leonard Edward**
Hatfield, PA 19440 (US)

• **Gaffney, Anne Mae**
West Chester, PA 19380 (US)
• **Han, Scott**
Lawrenceville, NJ 08648 (US)
• **Heffner, Michele Doreen**
Chalfont, PA 18914 (US)
• **Song, Ruozhi**
Wilmington, DE 19809 (US)

(74) Representative: **Kent, Venetia Katherine**
Rohm and Haas (UK) Ltd
European Operations Patent Dept.
Lennig House
2 Mason's Avenue
Croydon, CR9 3NB (GB)

(54) **Recalcined catalyst**

(57) A mixed metal oxide, which may be an orthorhombic phase material, is improved as a catalyst for the production of unsaturated carboxylic acids, or unsaturated nitriles, from alkanes, or mixtures of alkanes and alkenes, by: contacting with a liquid contacting member selected from the group consisting of organic

acids, alcohols, inorganic acids and hydrogen peroxide to form a contact mixture; recovering insoluble material from the contact mixture; and calcining the recovered insoluble material in a non-oxidizing atmosphere.

EP 1 254 709 A2

Description

[0001] The present invention relates to a catalyst for the oxidation of alkanes, or a mixture of alkanes and alkenes, to their corresponding unsaturated carboxylic acids by vapor phase catalytic oxidation and, more particularly, to a method of making the catalyst and to a process for the vapor phase catalytic oxidation of alkanes, or a mixture of alkanes and alkenes, to their corresponding unsaturated carboxylic acids using a catalyst prepared by the present method of making a catalyst. The present invention also relates to a process for the vapor phase catalytic oxidation of alkanes, or a mixture of alkanes and alkenes, in the presence of ammonia, to their corresponding unsaturated nitriles using a catalyst prepared by the present method of making a catalyst.

[0002] Nitriles, such as acrylonitrile and methacrylonitrile, have been industrially produced as important intermediates for the preparation of fibers, synthetic resins, synthetic rubbers, and the like. The most popular method for producing such nitriles is to subject an olefin such as propene or isobutene to a gas phase catalytic reaction with ammonia and oxygen in the presence of a catalyst at a high temperature. Known catalysts for conducting this reaction include a Mo-Bi-P-O catalyst, a V-Sb-O catalyst, an Sb-U-V-Ni-O catalyst, a Sb-Sn-O catalyst, a V-Sb-W-P-O catalyst and a catalyst obtained by mechanically mixing a V-Sb-W-O oxide and a Bi-Ce-Mo-W-O oxide. However, in view of the price difference between propane and propene or between isobutane and isobutene, attention has been drawn to the development of a method for producing acrylonitrile or methacrylonitrile by an ammoxidation reaction wherein a lower alkane, such as propane or isobutane, is used as a starting material, and it is catalytically reacted with ammonia and oxygen in a gaseous phase in the presence of a catalyst.

[0003] In particular, U.S. Patent No. 5,281,745 discloses a method for producing an unsaturated nitrile comprising subjecting an alkane and ammonia in the gaseous state to catalytic oxidation in the presence of a catalyst which satisfies the conditions:

(1) the mixed metal oxide catalyst is represented by the empirical formula



wherein X is at least one element selected from the group consisting of niobium, tantalum, tungsten, titanium, aluminum, zirconium, chromium, manganese, iron, ruthenium, cobalt, rhodium, nickel, palladium, platinum, antimony, bismuth, boron and cerium and, when $a = 1$, $b = 0.01$ to 1.0 , $c = 0.01$ to 1.0 , $x = 0.01$ to 1.0 and n is a number such that the total valency of the metal elements is satisfied; and

(2) the catalyst has X-ray diffraction peaks at the following angles ($\pm 0.3^\circ$) of 2θ in its X-ray diffraction pattern: 22.1° , 28.2° , 36.2° , 45.2° and 50.0° .

[0004] Similarly, Japanese Laid-Open Patent Application Publication No. 6-228073 discloses a method of nitrile preparation comprising reacting an alkane in a gas phase contact reaction with ammonia in the presence of a mixed metal oxide catalyst of the formula



wherein X represents one or more elements selected from niobium, tantalum, titanium, aluminum, zirconium, chromium, manganese, iron, ruthenium, cobalt, rhodium, nickel, palladium, platinum, antimony, bismuth, indium and cerium and, when $a = 1$, $b = 0.01$ to 1.0 , $c = 0.01$ to 1.0 , $x = 0.01$ to 1.0 and n is determined by the oxide form of the elements.

[0005] Unsaturated carboxylic acids such as acrylic acid and methacrylic acid are industrially important as starting materials for various synthetic resins, coating materials and plasticizers. Commercially, the current process for acrylic acid manufacture involves a two-step catalytic oxidation reaction starting with a propene feed. In the first stage, propene is converted to acrolein over a modified bismuth molybdate catalyst. In the second stage, acrolein product from the first stage is converted to acrylic acid using a catalyst composed of mainly molybdenum and vanadium oxides. In most cases, the catalyst formulations are proprietary to the catalyst supplier, but, the technology is well established. Moreover, there is an incentive to develop a single step process to prepare the unsaturated acid from its corresponding alkene. Therefore, the prior art describes cases where complex metal oxide catalysts are utilized for the preparation of unsaturated acid from a corresponding alkene in a single step.

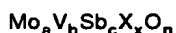
[0006] Japanese Laid-Open Patent Application Publication No. 07-053448 discloses the manufacture of acrylic acid by the gas-phase catalytic oxidation of propene in the presence of mixed metal oxides containing Mo, V, Te, O and X wherein X is at least one of Nb, Ta, W, Ti, Al, Zr, Cr, Mn, Fe, Ru, Co, Rh, Ni, Pd, Pt, Sb, Bi, B, In, Li, Na, K, Rb, Cs and Ce.

[0007] Commercial incentives also exist for producing acrylic acid using a lower cost propane feed. Therefore, the

prior art describes cases wherein a mixed metal oxide catalyst is used to convert propane to acrylic acid in one step. [0008] U.S. Patent No. 5,380,933 discloses a method for producing an unsaturated carboxylic acid comprising subjecting an alkane to a vapor phase catalytic oxidation reaction in the presence of a catalyst containing a mixed metal oxide comprising, as essential components, Mo, V, Te, O and X, wherein X is at least one element selected from the group consisting of niobium, tantalum, tungsten, titanium, aluminum, zirconium, chromium, manganese, iron, ruthenium, cobalt, rhodium, nickel, palladium, platinum, antimony, bismuth, boron, indium and cerium; and wherein the proportions of the respective essential components, based on the total amount of the essential components, exclusive of oxygen, satisfy the following relationships:

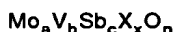
0.25 < r(Mo) < 0.98, 0.003 < r(V) < 0.5, 0.003 < r(Te) < 0.5 and 0.003 < r(X) < 0.5, wherein r(Mo), r(V), r(Te) and r(X) are the molar fractions of Mo, V, Te and X, respectively, based on the total amount of the essential components exclusive of oxygen.

[0009] Nonetheless, the prior art continues to seek ways to improve the performance of such mixed metal oxide catalysts. In particular, Japanese Laid-Open Patent Application Publication No. 10-330343 discloses a single stage washing or a dual stage washing followed by calcination of a mixed metal oxide of the formula



wherein X is at least one metal element selected from Ti, Zr, Nb, Ta, Cr, W, Mn, Fe, Ru, Co, Rh, Ir, Ni, Pd, Pt, Cu, Ag, Zn, In, Sn, Pb, Bi, Ce and alkaline earth metals, wherein, when $a=1$, $0.02 \leq b \leq 0.99$, $0.001 \leq c \leq 0.9$, $0 \leq x \leq 0.89$, $0.1 \leq c/b \leq 0.80$ and n is a value determined by the oxidation state of the other elements,

with solvents selected from aqueous oxalic acid, ethylene glycol or aqueous hydrogen peroxide. The so-formed catalyst is used for the ammoxidation of alkanes to form nitriles. Japanese Laid-Open Patent Application Publication No. 11-043314 discloses the washing of a mixed metal oxide of the formula



wherein X is at least one metal element selected from Ti, Zr, Nb, Ta, Cr, W, Mn, Fe, Ru, Co, Rh, Ir, Ni, Pd, Pt, Cu, Ag, Zn, In, Sn, Pb, Bi, Ce and alkaline earth metals,

wherein, when $a = 1$, $0.02 \leq b \leq 0.99$, $0.001 \leq c \leq 0.9$, $0 \leq x \leq 0.89$, $0.1 \leq c/b \leq 0.80$ and n is a value determined by the oxidation state of the other elements,

with at least one solvent selected from an aqueous solution of an organic acid, an alcohol, an aqueous solution of an inorganic acid or an aqueous solution of hydrogen peroxide, followed by calcination. The so-formed material is indicated to be useful in such applications as electronic materials, electrode materials, mechanical inorganic materials and as catalysts in petrochemistry, etc. In particular, use as a catalyst in the oxidative dehydrogenation of ethane to produce ethylene is exemplified.

[0010] It has now been found that the performance of a mixed metal oxide catalyst having the empirical formula



wherein

A is at least one element selected from the group consisting of Mo and W, N is at least one element selected from the group consisting of Te, Se and Sb, X is at least one element selected from the group consisting of Nb, Ta, Ti, Al, Zr, Cr, Mn, Fe, Ru, Co, Rh, Ni, Pt, Bi, B, In, Ce, As, Ge, Sn, Li, Na, K, Rb, Cs, Fr, Be, Mg, Ca, Sr, Ba, Ra, Hf, Pb, P, Pm, Eu, Gd, Dy, Ho, Er, Tm, Yb, Lu, Au, Ag, Re, Pr, Zn, Ga, Pd, Ir, Nd, Y, Sm, Tb, Br, Cu, Sc, Cl, F and I, wherein A, V, N and X are present in such amounts that the atomic ratio of A : V : N : X is $a : b : c : d$, and wherein, when $a = 1$, $b = 0.01$ to 2, $c = 0.01$ to 1.0, $d = 0.01$ to 1.0 and e is dependent on the oxidation state of the other elements,

may be improved for the oxidation (in the case where N is at least one element selected from the group consisting of Te, Se and Sb) of an alkane, or a mixture of an alkane and an alkene, to an unsaturated carboxylic acid or may be improved for the ammoxidation (in the case where N is at least one element selected from the group consisting of Te and Se) of an alkane, or a mixture of an alkane and an alkene, to an unsaturated nitrile, by contacting the mixed metal oxide catalyst with a liquid contact member selected from the group consisting of organic acids, alcohols, inorganic acids and hydrogen peroxide, followed by calcination of recovered insoluble materials. In fact, upon such contacting and subsequent calcination, both activity and selectivity may be improved, regardless of whether or not the mixed

metal oxide material is of a single crystalline phase or not.

[0011] Thus, in a first aspect, the present invention provides a process for preparing an improved catalyst, said process comprising:

(a) providing a mixed metal oxide having the empirical formula



wherein A is at least one element selected from the group consisting of Mo and W, N is at least one element selected from the group consisting of Te and Se, and X is at least one element selected from the group consisting of Nb, Ta, Ti, Al, Zr, Cr, Mn, Fe, Ru, Co, Rh, Ni, Pt, Bi, B, In, Ce, As, Ge, Sn, Li, Na, K, Rb, Cs, Fr, Be, Mg, Ca, Sr, Ba, Ra, Hf, Pb, P, Pm, Eu, Gd, Dy, Ho, Er, Tm, Yb, Lu, Au, Ag, Re, Pr, Zn, Ga, Pd, Ir, Nd, Y, Sm, Tb, Br, Cu, Sc, Cl, F and I,

wherein A, V, N and X are present in such amounts that the atomic ratio of A : V : N : X is a : b : c : d, and wherein, when a = 1, b = 0.1 to 2, c = 0.1 to 1, d = 0.01 to 1 and e is dependent on the oxidation state of the other elements;

(b) contacting said mixed metal oxide with a liquid contact member selected from the group consisting of organic acids, alcohols, inorganic acids and hydrogen peroxide to form a contact mixture;

(c) recovering insoluble material from said contact mixture; and

(d) calcining said recovered insoluble material to form said improved catalyst.

[0012] In a second aspect, the present invention provides a catalyst produced by the process according to the first aspect of the invention.

[0013] In a third aspect, the present invention provides a process for producing an unsaturated carboxylic acid which comprises subjecting an alkane, or a mixture of an alkane and an alkene, to a vapor phase catalytic oxidation reaction in the presence of a catalyst produced by the process comprising:

(a) providing a mixed metal oxide having the empirical formula



wherein A is at least one element selected from the group consisting of Mo and W, N is at least one element selected from the group consisting of Te, Se and Sb, and X is at least one element selected from the group consisting of Nb, Ta, Ti, Al, Zr, Cr, Mn, Fe, Ru, Co, Rh, Ni, Pt, Bi, B, In, Ce, As, Ge, Sn, Li, Na, K, Rb, Cs, Fr, Be, Mg, Ca, Sr, Ba, Ra, Hf, Pb, P, Pm, Eu, Gd, Dy, Ho, Er, Tm, Yb, Lu, Au, Ag, Re, Pr, Zn, Ga, Pd, Ir, Nd, Y, Sm, Tb, Br, Cu, Sc, Cl, F and I,

wherein A, V, N and X are present in such amounts that the atomic ratio of A : V : N : X is a : b : c : d, and wherein, when a = 1, b = 0.1 to 2, c = 0.1 to 1, d = 0.01 to 1 and e is dependent on the oxidation state of the other elements;

(b) contacting said mixed metal oxide with a liquid contact member selected from the group consisting of organic acids, alcohols, inorganic acids and hydrogen peroxide to form a contact mixture;

(c) recovering insoluble material from said contact mixture; and

(d) calcining said recovered insoluble material to form said catalyst.

[0014] In a fourth aspect, the present invention provides a process for producing an unsaturated nitrile which comprises subjecting an alkane, or a mixture of an alkane and an alkene, and ammonia to a vapor phase catalytic oxidation reaction in the presence of a catalyst produced by the process comprising:

(a) providing a mixed metal oxide having the empirical formula



wherein A is at least one element selected from the group consisting of Mo and W, N is at least one element selected from the group consisting of Te and Se, and X is at least one element selected from the group consisting of Nb, Ta, Ti, Al, Zr, Cr, Mn, Fe, Ru, Co, Rh, Ni, Pt, Bi, B, In, Ce, As, Ge, Sn, Li, Na, K, Rb, Cs, Fr, Be, Mg, Ca, Sr,

Ba, Ra, Hf, Pb, P, Pm, Eu, Gd, Dy, Ho, Er, Tm, Yb, Lu, Au, Ag, Re, Pr, Zn, Ga, Pd, Ir, Nd, Y, Sm, Tb, Br, Cu, Sc, Cl, F and I,

wherein A, V, N and X are present in such amounts that the atomic ratio of A : V : N : X is a : b : c : d, and

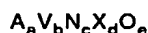
wherein, when a = 1, b = 0.1 to 2, c = 0.1 to 1, d = 0.01 to 1 and e is dependent on the oxidation state of the other elements;

(b) contacting said mixed metal oxide with a liquid contact member selected from the group consisting of organic acids, alcohols, inorganic acids and hydrogen peroxide to form a contact mixture;

(c) recovering insoluble material from said contact mixture; and

(d) calcining said recovered insoluble material to form said catalyst.

[0015] The mixed metal oxide, which is used as the starting material for the present process of preparing an improved catalyst has the empirical formula



wherein A is at least one element selected from the group consisting of Mo and W, N is at least one element selected from the group consisting of Te and Se, and X is at least one element selected from the group consisting of Nb, Ta, Ti, Al, Zr, Cr, Mn, Fe, Ru, Co, Rh, Ni, Pt, Bi, B, In, Ce, As, Ge, Sn, Li, Na, K, Rb, Cs, Fr, Be, Mg, Ca, Sr, Ba, Ra, Hf, Pb, P, Pm, Eu, Gd, Dy, Ho, Er, Tm, Yb, Lu, Au, Ag, Re, Pr, Zn, Ga, Pd, Ir, Nd, Y, Sm, Tb, Br, Cu, Sc, Cl, F and I; and wherein, when a = 1, b = 0.01 to 2, c = 0.01 to 1.0, d = 0.01 to 0.1 and e is dependent on the oxidation state of the other elements.

[0016] Preferably, when a = 1, b = 0.1 to 0.5, c = 0.05 to 0.5 and d = 0.01 to 0.5. More preferably, when a = 1, b = 0.15 to 0.45, c = 0.05 to 0.45 and d = 0.01 to 0.1. The value of e, i.e. the amount of oxygen present, is dependent on the oxidation state of the other elements in the catalyst. However, e is typically in the range of from 3 to 4.7.

[0017] For example, such a mixed metal oxide may be prepared by:

admixing compounds of elements A, V, N, X and at least one solvent to form a mixture,

wherein A is at least one element selected from the group consisting of Mo and W, N is at least one element selected from the group consisting of Te and Se, and X is at least one element selected from the group consisting of Nb, Ta, Ti, Al, Zr, Cr, Mn, Fe, Ru, Co, Rh, Ni, Pt, Bi, B, In, Ce, As, Ge, Sn, Li, Na, K, Rb, Cs, Fr, Be, Mg, Ca, Sr, Ba, Ra, Hf, Pb, P, Pm, Eu, Gd, Dy, Ho, Er, Tm, Yb, Lu, Au, Ag, Re, Pr, Zn, Ga, Pd, Ir, Nd, Y, Sm, Tb, Br, Cu, Sc, Cl, F and I,

wherein A, V, N and X are present in such amounts that the atomic ratio of A : V : N : X is a : b : c : d, and

wherein, when a = 1, b = 0.01 to 2, c = 0.01 to 1.0 and d = 0.01 to 1.0;

removing the at least one solvent from the mixture to form a precursor; and calcining the precursor to form a mixed metal oxide.

[0018] Preferred novel mixed metal oxides have the empirical formulae $Mo_aV_bTe_cNb_dO_e$ and $W_aV_bTe_cNb_dO_e$ wherein a, b, c, d and e are as previously defined.

[0019] Possibly, the mixed metal oxide to be used as the starting material is an orthorhombic phase material.

[0020] For example, the orthorhombic phase mixed metal oxide may be prepared by the process comprising:

(a) providing a mixed metal oxide having the empirical formula



wherein A is at least one element selected from the group consisting of Mo and W, N is at least one element selected from the group consisting of Te and Se, and X is at least one element selected from the group consisting of Nb, Ta, Ti, Al, Zr, Cr, Mn, Fe, Ru, Co, Rh, Ni, Pt, Bi, B, In, Ce, As, Ge, Sn, Li, Na, K, Rb, Cs, Fr, Be, Mg, Ca, Sr, Ba, Ra, Hf, Pb, P, Pm, Eu, Gd, Dy, Ho, Er, Tm, Yb, Lu, Au, Ag, Re, Pr, Zn, Ga, Pd, Ir, Nd, Y, Sm, Tb, Br, Cu, Sc, Cl, F and I,

wherein A, V, N and X are present in such amounts that the atomic ratio of A : V : N : X is a : b : c : d, and

wherein, when a = 1, b = 0.1 to 2, c = 0.1 to 1, d = 0.01 to 1 and e is dependent on the oxidation state of the other elements;

(b) contacting said mixed metal oxide with a liquid contact member selected from the group consisting of organic acids, alcohols, inorganic acids and hydrogen peroxide to form a contact mixture; and

(c) recovering insoluble material from said contact mixture to obtain said orthorhombic phase mixed metal oxide.

[0021] Alternatively, for example, such an orthorhombic phase mixed metal oxide may be prepared by a process comprising:

(a) admixing compounds of elements A, V, N and X and at least one solvent to form a solution,

wherein A is at least one element selected from the group consisting of Mo and W, N is at least one element selected from the group consisting of Te and Se, and X is at least one element selected from the group consisting of Nb, Ta, Ti, Al, Zr, Cr, Mn, Fe, Ru, Co, Rh, Ni, Pt, Bi, B, In, Ce, As, Ge, Sn, Li, Na, K, Rb, Cs, Fr, Be, Mg, Ca, Sr, Ba, Ra, Hf, Pb, P, Pm, Eu, Gd, Dy, Ho, Er, Tm, Yb, Lu, Au, Ag, Re, Pr, Zn, Ga, Pd, Ir, Nd, Y, Sm, Tb, Br, Cu, Sc, Cl, F and I,

wherein A, V, N and X are present in such amounts that the atomic ratio of A : V : N : X is a : b : c : d, and wherein, when a = 1, b = 0.01 to 2, c = 0.01 to 1.0 and d = 0.01 to 1.0;

(b) admixing a seeding effective amount of an orthorhombic phase mixed metal oxide seed, substantially free of hexagonal phase mixed metal oxide, with said solution to form a seeded solution;

(c) removing said at least one solvent from said seeded solution to form a catalyst precursor; and

(d) calcining said catalyst precursor to obtain said orthorhombic phase mixed metal oxide.

[0022] By a seeding effective amount is meant an amount of seed material effective to cause nucleation of the orthorhombic phase, e.g., 0.01 % by weight of seed material based on the total weight of the solution being seeded. By an orthorhombic phase mixed metal oxide which is substantially free of hexagonal phase mixed metal oxide is meant a material that contains not less than 90 % by weight of orthorhombic phase based on the total weight of the material.

[0023] The orthorhombic phase mixed metal oxide substantially free of hexagonal phase mixed metal oxide, to be used as seed, may be prepared, for example, by:

taking a mixed metal oxide having the empirical formula



wherein A is at least one element selected from the group consisting of Mo and W, N is at least one element selected from the group consisting of Te and Se, and X is at least one element selected from the group consisting of Nb, Ta, Ti, Al, Zr, Cr, Mn, Fe, Ru, Co, Rh, Ni, Pt, Bi, B, In, Ce, As, Ge, Sn, Li, Na, K, Rb, Cs, Fr, Be, Mg, Ca, Sr, Ba, Ra, Hf, Pb, P, Pm, Eu, Gd, Dy, Ho, Er, Tm, Yb, Lu, Au, Ag, Re, Pr, Zn, Ga, Pd, Ir, Nd, Y, Sm, Tb, Br, Cu, Sc, Cl, F and I,

wherein, when a = 1, b = 0.01 to 2, c = 0.01 to 1.0, d = 0.01 to 1.0 and e is dependent on the oxidation state of the other elements;

contacting the mixed metal oxide with a liquid contacting member selected from the group consisting of organic acids, alcohols, inorganic acids and hydrogen peroxide to form a contact mixture; and recovering insoluble material, to be used as seed material, from the contact mixture.

[0024] Alternatively, the orthorhombic phase mixed metal oxide substantially free of hexagonal phase mixed metal oxide, to be used as seed, may be prepared, for example, by:

admixing compounds of elements A, V, N and X and at least one solvent to form a mixture

wherein A is at least one element selected from the group consisting of Mo and W, N is at least one element selected from the group consisting of Te and Se, and X is at least one element selected from the group consisting of Nb, Ta, Ti, Al, Zr, Cr, Mn, Fe, Ru, Co, Rh, Ni, Pt, Bi, B, In, Ce, As, Ge, Sn, Li, Na, K, Rb, Cs, Fr, Be, Mg, Ca, Sr, Ba, Ra, Hf, Pb, P, Pm, Eu, Gd, Dy, Ho, Er, Tm, Yb, Lu, Au, Ag, Re, Pr, Zn, Ga, Pd, Ir, Nd, Y, Sm, Tb, Br, Cu, Sc, Cl, F and I,

wherein A, V, N and X are present in such amounts that the atomic ratio of A : V : N : X is a : b : c : d, and wherein, when a = 1, b = 0.01 to 2, c = 0.01 to 1.0 and d = 0.01 to 1.0;

removing the at least one solvent from the mixture to form a precursor;

calcining the precursor to form a calcined precursor;

contacting the calcined precursor with a liquid contact member selected from the group consisting of organic acids, alcohols, inorganic acids and hydrogen peroxide to form a contact mixture; and

recovering insoluble material, to be used as seed material, from the contact mixture.

[0025] Suitable solvents, for the above-noted processes, include water; alcohols including, but not limited to, meth-

anol, ethanol, propanol, and diols, etc.; as well as other polar solvents known in the art. Generally, water is preferred. The water is any water suitable for use in chemical syntheses including, without limitation, distilled water and de-ionized water. The amount of water present is preferably an amount sufficient to keep the elements substantially in solution long enough to avoid or minimize compositional and/or phase segregation during the preparation steps. Accordingly, the amount of water will vary according to the amounts and solubilities of the materials combined. However, the amount of water is preferably sufficient to ensure an aqueous solution is formed.

[0026] The solvent is removed by any suitable method, known in the art, to form a catalyst precursor. Such methods include, without limitation, vacuum drying, freeze drying, spray drying, rotary evaporation and air drying.

[0027] For example, in the case of water being the solvent: Vacuum drying is generally performed at pressures ranging from 10mmHg to 500mmHg. Freeze drying typically entails freezing the solution, using, for instance, liquid nitrogen, and drying the frozen solution under vacuum. Spray drying is generally performed under an inert atmosphere such as nitrogen or argon, with an inlet temperature ranging from 125°C to 200°C and an outlet temperature ranging from 75°C to 150°C. Rotary evaporation is generally performed at a bath temperature of from 25°C to 90°C and at a pressure of from 10mmHg to 760mmHg, preferably at a bath temperature of from 40° to 90°C and at a pressure of from 10mmHg to 350mmHg, more preferably at a bath temperature of from 40°C to 60°C and at a pressure of from 10mmHg to 40mmHg. Air drying may be effected at temperatures ranging from 25°C to 90°C. Rotary evaporation or spray drying are generally preferred.

[0028] Contacting with a liquid contact member selected from the group consisting of organic acids, alcohols, inorganic acids and hydrogen peroxide (whether it be for the purpose of making the final catalyst, an orthorhombic phase mixed metal oxide, or for the purpose of making a seed material) may be effected without any particular restrictions (so long as, in the case of the preparation of the orthorhombic phase mixed metal oxide, the hexagonal phase is fully removed from the mixed metal oxide or the calcined precursor). The liquid contact member is normally used in an amount of 1 to 100 times the volume of the mixed metal oxide or the first calcined precursor, preferably 3 to 50 times the volume, more preferably 5 to 25 times the volume. (Contacting at elevated temperatures will remove the hexagonal phase more rapidly. However, if prolonged contact time is not a consideration, contacting at room temperature may be utilized.) Normally, contact temperatures of room temperature to 100°C are utilized, preferably 50°C to 90°C, more preferably 60°C to 80°C. As previously noted, contact time will be affected by the temperature at which the contacting is carried out. Normally, contact times of 1 to 100 hours are utilized, preferably 2 to 20 hours, more preferably 5 to 10 hours. The contact mixture is preferably agitated during the contacting.

[0029] There are no particular restrictions upon the organic acids which may be used as the liquid contacting member. For example, oxalic acid, formic acid, acetic acid, citric acid and tartaric acid may be used, however, oxalic acid is preferred. If the organic acid is a liquid, it may be used as is or in an aqueous solution. If the organic acid is a solid, it is used in an aqueous solution. When using aqueous solutions, there are no particular restrictions on the concentration of the organic acid. Normally, the concentration of the organic acid in the aqueous solution can vary from 0.1 to 50% by weight, preferably 1 to 15 % by weight.

[0030] There are no particular restrictions upon the alcohols which may be used as the liquid contacting member. For example, methanol, ethanol, propanol, butanol, hexanol and diols may be utilized, however, alcohols having one to four carbon atoms are preferred, with ethylene glycol being particularly preferred. The alcohols may be utilized in the form of aqueous solutions, but, if so, the water content should be held to 20 % by weight or less for the best effectiveness.

[0031] Similarly, there are no particular restrictions upon the inorganic acids which may be used as the liquid contacting member. For example, telluric acid, nitric acid, sulfuric acid, phosphoric acid, hydrochloric acid, perchloric acid, chloric acid and hypochlorous acid may be used. The inorganic acids are typically used as aqueous solutions with concentrations of the acids in the range of from 0.1 to 50 % by weight, preferably from 0.1 to 10 % by weight.

[0032] When hydrogen peroxide is utilized as the liquid contacting member, it is used in the form of an aqueous solution having a concentration in the range of from 0.1 to 50 % by weight, preferably from 1 to 10% by weight.

[0033] After contacting with the liquid contacting member, insoluble material is recovered from the so-formed contact mixture. The insoluble material may be recovered by any conventional method, e.g., centrifugation or filtration. If the contacting was conducted at elevated temperature, the contact mixture may be cooled prior to recovery of the insoluble material. If desired, the recovered insoluble material may be washed with one of the previously disclosed solvents, preferably water.

[0034] Calcination may be conducted in an oxidizing atmosphere, e.g., in air, oxygen-enriched air or oxygen, or in a non-oxidizing atmosphere, e.g., in an inert atmosphere or in vacuo. The inert atmosphere may be any material which is substantially inert, i.e., does not react or interact with, the catalyst precursor. Suitable examples include, without limitation, nitrogen, argon, xenon, helium or mixtures thereof. Preferably, the inert atmosphere is argon or nitrogen. The inert atmosphere may flow over the surface of the catalyst precursor or may not flow thereover (a static environment). When the inert atmosphere does flow over the surface of the catalyst precursor, the flow rate can vary over a wide range, e.g., at a space velocity of from 1 to 500 hr⁻¹.

[0035] The calcination is usually performed at a temperature of from 350°C to 850°C, preferably from 400°C to 700°C, more preferably from 500°C to 640°C. The calcination is performed for an amount of time suitable to form the aforementioned catalyst. Typically, the calcination is performed for from 0.5 to 30 hours, preferably from 1 to 25 hours, more preferably for from 1 to 15 hours, to obtain the desired promoted mixed metal oxide.

[0036] In a preferred mode of operation, the catalyst precursor is calcined in two stages. In the first stage, the catalyst precursor is calcined in an oxidizing environment (e.g. air) at a temperature of from 200°C to 400°C, preferably from 275°C to 325°C for from 15 minutes to 8 hours, preferably for from 1 to 3 hours. In the second stage, the material from the first stage is calcined in a non-oxidizing environment (e.g., an inert atmosphere) at a temperature of from 500°C to 750°C, preferably for from 550°C to 650°C, for 15 minutes to 8 hours, preferably for from 1 to 3 hours. Optionally, a reducing gas, such as, for example, ammonia or hydrogen, may be added during the second stage calcination.

[0037] In a particularly preferred mode of operation, the catalyst precursor in the first stage is placed in the desired oxidizing atmosphere at room temperature and then raised to the first stage calcination temperature and held there for the desired first stage calcination time. The atmosphere is then replaced with the desired non-oxidizing atmosphere for the second stage calcination, the temperature is raised to the desired second stage calcination temperature and held there for the desired second stage calcination time.

[0038] In another preferred mode of operation, the insoluble material recovered from the contact mixture is calcined in a non-oxidizing atmosphere, preferably in an inert atmosphere.

[0039] Although any type of heating mechanism, e.g., a furnace, may be utilized during the calcination, it is preferred to conduct the calcination under a flow of the designated gaseous environment. Therefore, it is advantageous to conduct the calcination in a bed with continuous flow of the desired gas(es) through the bed of solid catalyst precursor particles.

[0040] The starting materials for the above mixed metal oxide are not limited to those described above. A wide range of materials including, for example, oxides, nitrates, halides or oxyhalides, alkoxides, acetylacetonates, and organometallic compounds may be used. For example, ammonium heptamolybdate may be utilized for the source of molybdenum in the catalyst. However, compounds such as MoO_3 , MoO_2 , MoCl_5 , MoOCl_4 , $\text{Mo}(\text{OC}_2\text{H}_5)_5$, molybdenum acetylacetonate, phosphomolybdic acid and silicomolybdic acid may also be utilized instead of ammonium heptamolybdate. Similarly, ammonium metavanadate may be utilized for the source of vanadium in the catalyst. However, compounds such as V_2O_5 , V_2O_3 , VOCl_3 , VCl_4 , $\text{VO}(\text{OC}_2\text{H}_5)_3$, vanadyl sulfate, vanadium acetylacetonate and vanadyl acetylacetonate may also be utilized instead of ammonium metavanadate. The tellurium source may include telluric acid, TeCl_4 , $\text{Te}(\text{OC}_2\text{H}_5)_5$, $\text{Te}(\text{OCH}(\text{CH}_3)_2)_4$ and TeO_2 . The niobium source may include ammonium niobium oxalate, Nb_2O_5 , NbCl_5 , niobic acid or $\text{Nb}(\text{OC}_2\text{H}_5)_5$ as well as the more conventional niobium oxalate.

[0041] A mixed metal oxide, thus obtained, exhibits excellent catalytic activities by itself. However, the mixed metal oxide may be converted to a catalyst having higher activities by grinding.

[0042] There is no particular restriction as to the grinding method, and conventional methods may be employed. As a dry grinding method, a method of using a gas stream grinder may, for example, be mentioned wherein coarse particles are permitted to collide with one another in a high speed gas stream for grinding. The grinding may be conducted not only mechanically but also by using a mortar or the like in the case of a small scale operation.

[0043] As a wet grinding method wherein grinding is conducted in a wet state by adding water or an organic solvent to the above mixed metal oxide, a conventional method of using a rotary cylinder-type medium mill or a medium-stirring type mill, may be mentioned. The rotary cylinder-type medium mill is a wet mill of the type wherein a container for the object to be ground is rotated, and it includes, for example, a ball mill and a rod mill. The medium-stirring type mill is a wet mill of the type wherein the object to be ground, contained in a container is stirred by a stirring apparatus, and it includes, for example, a rotary screw type mill, and a rotary disc type mill.

[0044] The conditions for grinding may suitably be set to meet the nature of the above-mentioned promoted mixed metal oxide, the viscosity, the concentration, etc. of the solvent used in the case of wet grinding, or the optimum conditions of the grinding apparatus.

[0045] Further, in some cases, it is possible to further improve the catalytic activities by further adding a solvent to the ground catalyst precursor to form a solution or slurry, followed by drying again. There is no particular restriction as to the concentration of the solution or slurry, and it is usual to adjust the solution or slurry so that the total amount of the starting material compounds for the ground catalyst precursor is from 10 to 60 wt %. Then, this solution or slurry is dried by a method such as spray drying, freeze drying, evaporation to dryness or vacuum drying, preferably by the spray drying method. Further, similar drying may be conducted also in the case where wet grinding is conducted.

[0046] The oxide obtained by the above-mentioned method may be used as a final catalyst, but it may further be subjected to heat treatment usually at a temperature of from 200° to 700°C for from 0.1 to 10 hours.

[0047] The mixed metal oxide thus obtained may be used by itself as a solid catalyst, but may be formed into a catalyst together with a suitable carrier such as silica, alumina, titania, aluminosilicate, diatomaceous earth or zirconia. Further, it may be molded into a suitable shape and particle size depending upon the scale or system of the reactor.

[0048] Alternatively, the metal components of the presently contemplated catalyst may be supported on materials such as alumina, silica, silica-alumina, zirconia, titania, etc. by conventional incipient wetness techniques. In one typical

method, solutions containing the metals, whether seeded with the orthorhombic phase mixed metal oxide material or not, are contacted with the dry support such that the support is wetted; then, the resultant wetted material is dried, for example, at a temperature from room temperature to 200°C followed by calcination, as described above, liquid contacting and then re-calcination.

[0049] In a third aspect, the present invention provides a process for producing an unsaturated carboxylic acid, which comprises subjecting an alkane, or a mixture of an alkane and an alkene, to a vapor phase catalytic oxidation reaction in the presence of a catalyst produced in accord with the present invention to produce an unsaturated carboxylic acid.

[0050] In the production of such an unsaturated carboxylic acid, it is preferred to employ a starting material gas which contains steam. In such a case, as a starting material gas to be supplied to the reaction system, a gas mixture comprising a steam-containing alkane, or a steam-containing mixture of alkane and alkene, and an oxygen-containing gas, is usually used. However, the steam-containing alkane, or the steam-containing mixture of alkane and alkene, and the oxygen-containing gas may be alternately supplied to the reaction system. The steam to be employed may be present in the form of steam gas in the reaction system, and the manner of its introduction is not particularly limited.

[0051] Further, as a diluting gas, an inert gas such as nitrogen, argon or helium may be supplied. The molar ratio (alkane or mixture of alkane and alkene) : (oxygen) : (diluting gas) : (H₂O) in the starting material gas is preferably (1) : (0.1 to 10) : (0 to 20) : (0.2 to 70), more preferably (1) : (1 to 5.0) : (0 to 10) : (5 to 40).

[0052] When steam is supplied together with the alkane, or the mixture of alkane and alkene, as starting material gas, the selectivity for an unsaturated carboxylic acid is distinctly improved, and the unsaturated carboxylic acid can be obtained from the alkane, or mixture of alkane and alkene, in good yield simply by contacting in one stage. However, the conventional technique utilizes a diluting gas such as nitrogen, argon or helium for the purpose of diluting the starting material. As such a diluting gas, to adjust the space velocity, the oxygen partial pressure and the steam partial pressure, an inert gas such as nitrogen, argon or helium may be used together with the steam.

[0053] As the starting material alkane it is preferred to employ a C₃₋₈alkane, particularly propane, isobutane or n-butane; more preferably, propane or isobutane; most preferably, propane. According to the present invention, from such an alkane, an unsaturated carboxylic acid such as an α,β -unsaturated carboxylic acid can be obtained in good yield. For example, when propane or isobutane is used as the starting material alkane, acrylic acid or methacrylic acid will be obtained, respectively, in good yield.

[0054] In the present invention, as the starting material mixture of alkane and alkene, it is possible to employ a mixture of C₃₋₈alkane and C₃₋₈alkene, particularly propane and propene, isobutane and isobutene or n-butane and n-butene. As the starting material mixture of alkane and alkene, propane and propene or isobutane and isobutene are more preferred. Most preferred is a mixture of propane and propene. According to the present invention, from such a mixture of an alkane and an alkene, an unsaturated carboxylic acid such as an α,β -unsaturated carboxylic acid can be obtained in good yield. For example, when propane and propene or isobutane and isobutene are used as the starting material mixture of alkane and alkene, acrylic acid or methacrylic acid will be obtained, respectively, in good yield. Preferably, in the mixture of alkane and alkene, the alkene is present in an amount of at least 0.5% by weight, more preferably at least 1.0% by weight to 95% by weight; most preferably, 3% by weight to 90% by weight.

[0055] As an alternative, an alkanol, such as isobutanol, which will dehydrate under the reaction conditions to form its corresponding alkene, i.e. isobutene, may also be used as a feed to the present process or in conjunction with the previously mentioned feed streams.

[0056] The purity of the starting material alkane is not particularly limited, and an alkane containing a lower alkane such as methane or ethane, air or carbon dioxide, as impurities, may be used without any particular problem. Further, the starting material alkane may be a mixture of various alkanes. Similarly, the purity of the starting material mixture of alkane and alkene is not particularly limited, and a mixture of alkane and alkene containing a lower alkene such as ethene, a lower alkane such as methane or ethane, air or carbon dioxide, as impurities, may be used without any particular problem. Further, the starting material mixture of alkane and alkene may be a mixture of various alkanes and alkenes.

[0057] There is no limitation on the source of the alkene. It may be purchased, per se, or in admixture with an alkane and/or other impurities. Alternatively, it can be obtained as a byproduct of alkane oxidation. Similarly, there is no limitation on the source of the alkane. It may be purchased, per se, or in admixture with an alkene and/or other impurities. Moreover, the alkane, regardless of source, and the alkene, regardless of source, may be blended as desired.

[0058] The detailed mechanism of the oxidation reaction of the present invention is not clearly understood, but the oxidation reaction is carried out by oxygen atoms present in the above promoted mixed metal oxide or by molecular oxygen present in the feed gas. To incorporate molecular oxygen into the feed gas, such molecular oxygen may be pure oxygen gas. However, it is usually more economical to use an oxygen-containing gas such as air, since purity is not particularly required.

[0059] It is also possible to use only an alkane, or a mixture of alkane and alkene, substantially in the absence of molecular oxygen for the vapor phase catalytic reaction. In such a case, it is preferred to adopt a method wherein a part of the catalyst is appropriately withdrawn from the reaction zone from time to time, then sent to an oxidation

regenerator, regenerated and then returned to the reaction zone for reuse. As the regeneration method of the catalyst, a method may, for example, be mentioned which comprises contacting an oxidative gas such as oxygen, air or nitrogen monoxide with the catalyst in the regenerator usually at a temperature of from 300° to 600°C.

[0060] This third aspect of the present invention will be described in further detail with respect to a case where propane is used as the starting material alkane and air is used as the oxygen source. The reaction system may be a fixed bed system or a fluidized bed system. However, since the reaction is an exothermic reaction, a fluidized bed system may preferably be employed whereby it is easy to control the reaction temperature. The proportion of air to be supplied to the reaction system is important for the selectivity for the resulting acrylic acid, and it is usually at most 25 moles, preferably from 0.2 to 18 moles per mole of propane, whereby high selectivity for acrylic acid can be obtained. This reaction can be conducted usually under atmospheric pressure, but may be conducted under a slightly elevated pressure or slightly reduced pressure. With respect to other alkanes such as isobutane, or to mixtures of alkanes and alkenes such as propane and propene, the composition of the feed gas may be selected in accordance with the conditions for propane.

[0061] Typical reaction conditions for the oxidation of propane or isobutane to acrylic acid or methacrylic acid may be utilized in the practice of the present invention. The process may be practiced in a single pass mode (only fresh feed is fed to the reactor) or in a recycle mode (at least a portion of the reactor effluent is returned to the reactor). General conditions for the process of the present invention are as follows: the reaction temperature can vary from 200°C to 700°C, but is usually in the range of from 200°C to 550°C, more preferably 250°C to 480°C, most preferably 300°C to 400°C; the gas space velocity, SV, in the vapor phase reaction is usually within a range of from 100 to 10,000 hr⁻¹, preferably 300 to 6,000 hr⁻¹, more preferably 300 to 2,000 hr⁻¹; the average contact time with the catalyst can be from 0.01 to 10 seconds or more, but is usually in the range of from 0.1 to 10 seconds, preferably from 2 to 6 seconds; the pressure in the reaction zone usually ranges from 0 to 75 psig, but is preferably no more than 50 psig. In a single pass mode process, it is preferred that the oxygen be supplied from an oxygen-containing gas such as air. The single pass mode process may also be practiced with oxygen addition. In the practice of the recycle mode process, oxygen gas by itself is the preferred source so as to avoid the build up of inert gases in the reaction zone.

[0062] Of course, in the oxidation reaction of the present invention, it is important that the hydrocarbon and oxygen concentrations in the feed gases be maintained at the appropriate levels to minimize or avoid entering a flammable regime within the reaction zone or especially at the outlet of the reactor zone. Generally, it is preferred that the outlet oxygen levels be low to both minimize after-burning and, particularly, in the recycle mode of operation, to minimize the amount of oxygen in the recycled gaseous effluent stream. In addition, operation of the reaction at a low temperature (below 450°C) is extremely attractive because after-burning becomes less of a problem which enables the attainment of higher selectivity to the desired products. The catalyst of the present invention operates more efficiently at the lower temperature range set forth above, significantly reducing the formation of acetic acid and carbon oxides, and increasing selectivity to acrylic acid. As a diluting gas to adjust the space velocity and the oxygen partial pressure, an inert gas such as nitrogen, argon or helium may be employed.

[0063] When the oxidation reaction of propane, and especially the oxidation reaction of propane and propene, is conducted by the method of the present invention, carbon monoxide, carbon dioxide, acetic acid, etc. may be produced as by-products, in addition to acrylic acid. Further, in the method of the present invention, an unsaturated aldehyde may sometimes be formed depending upon the reaction conditions. For example, when propane is present in the starting material mixture, acrolein may be formed; and when isobutane is present in the starting material mixture, methacrolein may be formed. In such a case, such an unsaturated aldehyde can be converted to the desired unsaturated carboxylic acid by subjecting it again to the vapor phase catalytic oxidation with the promoted mixed metal oxide-containing catalyst of the present invention or by subjecting it to a vapor phase catalytic oxidation reaction with a conventional oxidation reaction catalyst for an unsaturated aldehyde.

[0064] In a fourth aspect, the present invention provides a process for producing an unsaturated nitrile, which comprises subjecting an alkane, or a mixture of an alkane and an alkene, to a vapor phase catalytic oxidation reaction with ammonia in the presence of a catalyst produced in accord with the present invention to produce an unsaturated nitrile.

[0065] In the production of such an unsaturated nitrile, as the starting material alkane, it is preferred to employ a C₃₋₈alkane such as propane, butane, isobutane, pentane, hexane and heptane. However, in view of the industrial application of nitriles to be produced, it is preferred to employ a lower alkane having 3 or 4 carbon atoms, particularly propane and isobutane.

[0066] Similarly, as the starting material mixture of alkane and alkene, it is possible to employ a mixture of C₃₋₈alkane and C₃₋₈alkene such as propane and propene, butane and butene, isobutane and isobutene, pentane and pentene, hexane and hexene, and heptane and heptene. However, in view of the industrial application of nitriles to be produced, it is more preferred to employ a mixture of a lower alkane having 3 or 4 carbon atoms and a lower alkene having 3 or 4 carbon atoms, particularly propane and propene or isobutane and isobutene. Preferably, in the mixture of alkane and alkene, the alkene is present in an amount of at least 0.5% by weight, more preferably at least 1.0% by weight to 95% by weight, most preferably 3% by weight to 90% by weight.

[0067] The purity of the starting material alkane is not particularly limited, and an alkane containing a lower alkane such as methane or ethane, air or carbon dioxide, as impurities, may be used without any particular problem. Further, the starting material alkane may be a mixture of various alkanes. Similarly, the purity of the starting material mixture of alkane and alkene is not particularly limited, and a mixture of alkane and alkene containing a lower alkene such as ethene, a lower alkane such as methane or ethane, air or carbon dioxide, as impurities, may be used without any particular problem. Further, the starting material mixture of alkane and alkene may be a mixture of various alkanes and alkenes.

[0068] There is no limitation on the source of the alkene. It may be purchased, per se, or in admixture with an alkane and/or other impurities. Alternatively, it can be obtained as a byproduct of alkane oxidation. Similarly, there is no limitation on the source of the alkane. It may be purchased, per se, or in admixture with an alkene and/or other impurities. Moreover, the alkane, regardless of source, and the alkene, regardless of source, may be blended as desired.

[0069] The detailed mechanism of the ammoxidation reaction of this aspect of the present invention is not clearly understood. However, the oxidation reaction is conducted by the oxygen atoms present in the above promoted mixed metal oxide or by the molecular oxygen in the feed gas. When molecular oxygen is incorporated in the feed gas, the oxygen may be pure oxygen gas. However, since high purity is not required, it is usually economical to use an oxygen-containing gas such as air.

[0070] As the feed gas, it is possible to use a gas mixture comprising an alkane, or a mixture of an alkane and an alkene, ammonia and an oxygen-containing gas. However, a gas mixture comprising an alkane or a mixture of an alkane and an alkene and ammonia, and an oxygen-containing gas may be supplied alternately.

[0071] When the gas phase catalytic reaction is conducted using an alkane, or a mixture of an alkane and an alkene, and ammonia substantially free from molecular oxygen, as the feed gas, it is advisable to employ a method wherein a part of the catalyst is periodically withdrawn and sent to an oxidation regenerator for regeneration, and the regenerated catalyst is returned to the reaction zone. As a method for regenerating the catalyst, a method may be mentioned wherein an oxidizing gas such as oxygen, air or nitrogen monoxide is permitted to flow through the catalyst in the regenerator usually at a temperature of from 300°C to 600°C.

[0072] This fourth aspect of the present invention will be described in further detail with respect to a case where propane is used as the starting material alkane and air is used as the oxygen source. The proportion of air to be supplied for the reaction is important with respect to the selectivity for the resulting acrylonitrile. Namely, high selectivity for acrylonitrile is obtained when air is supplied within a range of at most 25 moles, particularly 1 to 15 moles, per mole of the propane. The proportion of ammonia to be supplied for the reaction is preferably within a range of from 0.2 to 5 moles, particularly from 0.5 to 3 moles, per mole of propane. This reaction may usually be conducted under atmospheric pressure, but may be conducted under a slightly increased pressure or a slightly reduced pressure. With respect to other alkanes such as isobutane, or to mixtures of alkanes and alkenes such as propane and propene, the composition of the feed gas may be selected in accordance with the conditions for propane.

[0073] The process of this fourth aspect of the present invention may be conducted at a temperature of, for example, from 250°C to 500°C. More preferably, the temperature is from 300°C to 460°C. The gas space velocity, SV, in the gas phase reaction is usually within the range of from 100 to 10,000 hr⁻¹, preferably from 300 to 6,000 hr⁻¹, more preferably from 300 to 2,000 hr⁻¹. As a diluent gas, for adjusting the space velocity and the oxygen partial pressure, an inert gas such as nitrogen, argon or helium can be employed. When ammoxidation of propane is conducted by the method of the present invention, in addition to acrylonitrile, carbon monoxide, carbon dioxide, acetonitrile, hydrocyanic acid and acrolein may form as by-products.

Examples

Preparative

Example 1

[0074] Ammonium heptamolybdate (51.35 g), ammonium metavanadate (10.11 g) and telluric acid (15.36 g) were dissolved in warm water. Oxalic acid dihydrate (5.68 g) was dissolved in a 6.5% w/w solution of niobium oxalate in water (249.55 g), and the resulting niobium oxalate solution was added to the first solution. After stirring 15 minutes, the mixture was dried in a rotary evaporator, and then overnight under vacuum. The dried solid was calcined 1 hour at 275°C in air, and then 2 hours at 600°C in argon.

Example 2

[0075] The calcined solid of Example 1 (30.93 g) was added to a solution of oxalic acid dihydrate (12.38 g) in water (130 g). The mixture was heated with stirring 6 hours at 75°C, after which the mixture was cooled to room temperature

and filtered to collect insolubles. The solid was dried under vacuum 16 hours at room temperature.

Example 3

5 [0076] The dried solid of Example 2 was calcined 2 hours at 600°C in argon.

Example 4

10 [0077] The calcined solid of Example 3 was ground in a porcelain mortar and pestle.

Example 5

15 [0078] The calcined solid of Example 1 (11.38 g) was added to a solution of telluric acid (8.5 g) in water (60 g). The mixture was heated with stirring 6 hours at 75°C, after which the mixture was cooled to room temperature and filtered to collect insolubles. The solid was dried under vacuum 16 hours at room temperature.

Example 6

20 [0079] The dried solid of Example 5 was calcined 2 hours at 600°C in argon.

Example 7

[0080] The calcined solid of Example 6 was ground in a porcelain mortar and pestle.

25 Example 8

[0081] Ammonium heptamolybdate (36.77 g), ammonium metavanadate (7.30 g) and telluric acid (11.00 g) were dissolved in warm water. A 6.5 % w/w solution of niobium oxalate in water (160.91 g) was added to the first solution. The mixture was dried in a rotary evaporator, and then overnight under vacuum. The dried solid was calcined 1 hour at 200°C in air, and then 2 hours at 600°C in argon.

30 [0082] A portion of this material (14.3 g) was added to a solution of oxalic acid dihydrate (7.15 g) in water (135 g). The mixture was heated with stirring 6 hours at 70°C, after which the mixture was filtered and the solid fraction washed with water. The solid was dried under vacuum 16 hours at room temperature to yield material A.

35 [0083] Ammonium heptamolybdate (51.35 g), ammonium metavanadate (10.12 g) and telluric acid (15.36 g) were dissolved in warm water. Oxalic acid dihydrate (5.68 g) was dissolved in a 6.5 % w/w solution of niobium oxalate in water (249.55 g), and the resulting niobium oxalate solution was added to the first solution along with material A (0.30 g). After stirring 15 minutes, the mixture was dried in a rotary evaporator, and then overnight under vacuum. The dried solid was calcined 1 hour at 275°C in air, and then 2 hours at 600°C in argon.

40 Example 9

[0084] The calcined solid of Example 8 (25 g) was added to a solution of oxalic acid dihydrate (10 g) in water (100 g). The mixture was heated with stirring 6 hours at 75°C, then stirred 16 hours at room temperature. The mixture was filtered, and the solid was washed with water and dried under vacuum 16 hours at room temperature.

45 Example 10

[0085] The dried solid of Example 9 was calcined 2 hours at 600°C in argon.

50 Example 11

[0086] The calcined solid of Example 10 was ground in a porcelain mortar and pestle.

Example 12

55

[0087] The calcined solid of Example 8 (13.25 g) was added to a solution of telluric acid (9.66 g) in water (95 g). The mixture was heated with stirring 6 hours at 75°C, then stirred 16 hours at room temperature. The mixture was filtered, and the solid was washed with water and dried under vacuum 16 hours at room temperature.

Example 13

[0088] The dried solid of Example 12 was calcined 2 hours at 600°C in argon.

Comparative Example 1

[0089] 200 mL of an aqueous solution containing ammonium heptamolybdate tetrahydrate (1.0M Mo), ammonium metavanadate (0.3M V) and telluric acid (0.23M Te) formed by dissolving the corresponding salts in water at 70°C, was added to a 1000mL rotavap flask. Then 100 mL of an aqueous solution of niobium oxalate (0.16M Nb) and oxalic acid (0.58M) were added thereto. After removing the water via a rotary evaporator with a warm water bath at 50°C and 28mm Hg, the solid materials were further dried in a vacuum oven at 25°C overnight and then calcined. (Calcination was effected by placing the solid materials in an air atmosphere and then heating them to 275°C at 10°C/min and holding them under the air atmosphere at 275°C for one hour; the atmosphere was then changed to argon and the material was heated from 275°C to 600°C at 2°C/min and the material was held under the argon atmosphere at 600°C for two hours.) The final catalyst had a nominal composition of $\text{Mo}_1\text{V}_{0.3}\text{Te}_{0.23}\text{Nb}_{0.08}\text{O}_e$. The catalyst, thus obtained, was pressed in a mold and then broken and sieved to 10 - 20 mesh granules for reactor evaluation.

Example 14

[0090] 20 grams of the catalyst from Comparative Example 1 were ground and added to 400mL of 5 % aqueous HNO_3 . The resulting suspension was stirred at 25°C for 30 minutes, then the solids were collected by gravity filtration, dried in a vacuum oven overnight at 25°C, and heated in argon from 25°C to 600°C at 10°C/min and held at 600°C for two hours. The catalyst, thus obtained, was pressed in a mold and then broken and sieved to 10 - 20 mesh granules for reactor evaluation.

Evaluation and Results

Examples 2 - 7 and 9 - 13

[0091] The catalysts were evaluated in the partial oxidation of propane to acrylic acid. The feed composition for the partial oxidation was 1.0 mol % propane in air, which had been saturated with water at ambient conditions (i.e. 2.9 - 3.1 mol %). The feed composition flow rate was 11.5 cc/min, providing a residence time (based on catalyst volume) of 3 seconds.

Table 1

Example	Temperature (°C)	Propane Conversion (%)	Selectivity to Acrylic Acid (%)	Acrylic Acid Yield (%)
2	373	36	64	23
3	384	48	65	32
4	371	49	69	34
5	385	23	57	13
6	394	26	92	24
7	385	45	73	33
9	359	44	66	29
10	381	39	82	32
11	391	52	67	35
12	381	29	55	16
13	391	50	70	35

Comparative Example 1 and Example 14

[0092] Catalysts were evaluated in a 10 cm long Pyrex tube reactor (internal diameter: 3.9 mm). The catalyst bed

(4 cm long) was positioned with glass wool at approximately mid-length in the reactor and was heated with an electric furnace. Mass flow controllers and meters regulated the gas flow rate. The oxidation was conducted using a feed gas stream of propane, steam and air, with a feed ratio of propane:steam:air of 1:3:96. The reactor effluent was analyzed by an FTIR. The results at 390°C and at a 3 second residence time are shown in Table 2.

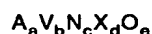
Table 2

	Propane Conversion (%)	Selectivity to Acrylic Acid (%)	Acrylic Acid Yield (%)
Comp. Example 1	41	41	17
Example 14	65	77	50

Claims

1. A process for preparing an improved catalyst, said process comprising:

(a) providing a mixed metal oxide having the empirical formula



wherein A is at least one element selected from the group consisting of Mo and W, N is at least one element selected from the group consisting of Te and Se, and X is at least one element selected from the group consisting of Nb, Ta, Ti, Al, Zr, Cr, Mn, Fe, Ru, Co, Rh, Ni, Pt, Bi, B, In, Ce, As, Ge, Sn, Li, Na, K, Rb, Cs, Fr, Be, Mg, Ca, Sr, Ba, Ra, Hf, Pb, P, Pm, Eu, Gd, Dy, Ho, Er, Tm, Yb, Lu, Au, Ag, Re, Pr, Zn, Ga, Pd, Ir, Nd, Y, Sm, Tb, Br, Cu, Sc, Cl, F and I,

wherein A, V, N and X are present in such amounts that the atomic ratio of A : V : N : X is a : b : c : d, and

wherein, when a = 1, b = 0.1 to 2, c = 0.1 to 1, d = 0.01 to 1 and e is dependent on the oxidation state of the other elements;

(b) contacting said mixed metal oxide with a liquid contact member selected from the group consisting of organic acids, alcohols, inorganic acids and hydrogen peroxide to form a contact mixture;

(c) recovering insoluble material from said contact mixture; and

(d) calcining said recovered insoluble material in a non-oxidizing atmosphere to form said improved catalyst.

2. The process for preparing an improved catalyst according to claim 1, wherein the mixed metal oxide of (a) is an orthorhombic phase mixed metal oxide.
3. The process for preparing an improved catalyst according to claim 1, wherein said liquid contact member is an aqueous solution of oxalic acid.
4. The process for preparing an improved catalyst according to claim 1, wherein said liquid contact member is an aqueous solution of telluric acid.
5. The process for preparing an improved catalyst according to claim 1, wherein said liquid contact member is an aqueous solution of nitric acid.
6. The process for preparing an improved catalyst according to claim 1, wherein said calcined recovered solid material is ground.
7. A catalyst produced by the process according to claim 1.
8. A process for producing an unsaturated nitrile which comprises subjecting an alkane, or a mixture of an alkane and an alkene, and ammonia to a vapor phase catalytic oxidation reaction in the presence of a catalyst produced by the process according to claim 1.
9. A process for producing an unsaturated carboxylic acid which comprises subjecting an alkane, or a mixture of an alkane and an alkene, to a vapor phase catalytic oxidation reaction in the presence of a catalyst produced by the

process comprising:

(a) providing a mixed metal oxide having the empirical formula



wherein A is at least one element selected from the group consisting of Mo and W, N is at least one element selected from the group consisting of Te, Se and Sb, and X is at least one element selected from the group consisting of Nb, Ta, Ti, Al, Zr, Cr, Mn, Fe, Ru, Co, Rh, Ni, Pt, Bi, B, In, Ce, As, Ge, Sn, Li, Na, K, Rb, Cs, Fr, Be, Mg, Ca, Sr, Ba, Ra, Hf, Ag, Pb, P, Pm, Eu, Gd, Dy, Ho, Er, Tm, Yb, Lu, Au, Ag, Re, Pr, Zn, Ga, Pd, Ir, Nd, Y, Sm, Tb, Br, Cu, Sc, Cl, F and I,

wherein A, V, N and X are present in such amounts that the atomic ratio of A : V : N : X is a : b : c : d, and wherein, when a = 1, b = 0.1 to 2, c = 0.1 to 1, d = 0.01 to 1 and e is dependent on the oxidation state of the other elements,

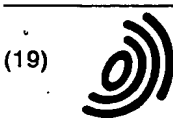
(b) contacting said mixed metal oxide with a liquid contact member selected from the group consisting of organic acids, alcohols, inorganic acids and hydrogen peroxide to form a contact mixture;

(c) recovering insoluble material from said contact mixture; and

(d) calcining said recovered insoluble material in a non-oxidizing atmosphere to form said catalyst.

10. The process for producing an unsaturated carboxylic acid according to claim 9, wherein the mixed metal oxide of (a) is an orthorhombic phase mixed metal oxide.

11. The process for producing an unsaturated carboxylic acid according to claim 8, wherein the calcined recovered material is ground.



Europäisches Patentamt
European Patent Office
Office européen des brevets



(11) **EP 1 254 709 A3**

(12) **EUROPEAN PATENT APPLICATION**

(88) Date of publication A3:
02.04.2003 Bulletin 2003/14

(51) Int Cl.7: **B01J 23/00, B01J 23/20,
B01J 23/28, C07C 253/24,
C07C 51/215**

(43) Date of publication A2:
06.11.2002 Bulletin 2002/45

(21) Application number: 02252866.5

(22) Date of filing: 23.04.2002

(84) Designated Contracting States:
**AT BE CH CY DE DK ES FI FR GB GR IE IT LI LU
MC NL PT SE TR**
Designated Extension States:
AL LT LV MK RO SI

- **Gaffney, Anne Mae**
West Chester, PA 19380 (US)
- **Han, Scott**
Lawrenceville, NJ 08648 (US)
- **Heffner, Michele Doreen**
Chalfont, PA 18914 (US)
- **Song, Ruozhi**
Wilmington, DE 19809 (US)

(30) Priority: 25.04.2001 US 286217 P

(71) Applicant: **ROHM AND HAAS COMPANY**
Philadelphia, Pennsylvania 19106-2399 (US)

(74) Representative: **Kent, Venetia Katherine**
Rohm and Haas (UK) Ltd
European Operations Patent Dept.
Lennig House
2 Mason's Avenue
Croydon, CR9 3NB (GB)

(72) Inventors:
• **Bogan Jr., Leonard Edward**
Hatfield, PA 19440 (US)

(54) **Recalcined catalyst**

(57) A mixed metal oxide, which may be an orthorhombic phase material, is improved as a catalyst for the production of unsaturated carboxylic acids, or unsaturated nitriles, from alkanes, or mixtures of alkanes and alkenes, by: contacting with a liquid contacting member selected from the group consisting of organic

acids, alcohols, inorganic acids and hydrogen peroxide to form a contact mixture; recovering insoluble material from the contact mixture; and calcining the recovered insoluble material in a non-oxidizing atmosphere.

EP 1 254 709 A3



European Patent
Office

PARTIAL EUROPEAN SEARCH REPORT

Application Number

which under Rule 45 of the European Patent Convention EP 02 25 2866
shall be considered, for the purposes of subsequent
proceedings, as the European search report

DOCUMENTS CONSIDERED TO BE RELEVANT			
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	CLASSIFICATION OF THE APPLICATION (Int.Cl.7)
X	US 6 063 728 A (KOMADA SATORU ET AL) 16 May 2000 (2000-05-16) * example 1 *	1-4,6-11	B01J23/00 B01J23/20 B01J23/28 C07C253/24 C07C51/215
E	EP 1 254 707 A (ROHM & HAAS) 6 November 2002 (2002-11-06) * page 7, paragraph 27 - page 8, paragraph 36; claims *	1-11	
P,X	EP 1 192 985 A (ROHM & HAAS) 3 April 2002 (2002-04-03) * the whole document *	1-4,6-11	
X	EP 0 962 253 A (ROHM & HAAS) 8 December 1999 (1999-12-08) * claims 1,5,9,12 *	1,2,4,6, 7,9-11	
D,A	EP 0 529 853 A (MITSUBISHI CHEM IND) 3 March 1993 (1993-03-03) * page 2, line 2-5 * * page 3, line 57 - page 4, line 17 * --- -/--	1-11	
			TECHNICAL FIELDS SEARCHED (Int.Cl.7)
			B01J C07C
INCOMPLETE SEARCH			
<p>The Search Division considers that the present application, or one or more of its claims, does/do not comply with the EPC to such an extent that a meaningful search into the state of the art cannot be carried out, or can only be carried out partially, for these claims.</p> <p>Claims searched completely :</p> <p>Claims searched incompletely</p> <p>Claims not searched :</p> <p>Reason for the limitation of the search:</p> <p>see sheet C</p>			
Place of search		Date of completion of the search	Examiner
BERLIN		11 February 2003	Gruber, M
CATEGORY OF CITED DOCUMENTS		<p>T : theory or principle underlying the invention</p> <p>E : earlier patent document, but published on, or after the filing date</p> <p>D : document cited in the application</p> <p>I : document cited for other reasons</p> <p>X : member of the same patent family, corresponding document</p>	
<p>X : particularly relevant if taken alone</p> <p>Y : particularly relevant if combined with another document of the same category</p> <p>A : technological background</p> <p>O : non-written disclosure</p> <p>P : intermediate document</p>			

EPC FORM 1503 03.02 (P04007)

European Patent
OfficeINCOMPLETE SEARCH
SHEET CApplication Number
EP 02 25 2866

Claim(s) searched incompletely:
1-11

Reason for the limitation of the search:

Present claims 1-11 relate to an extremely large number of possible compounds and methods. Support within the meaning of Article 84 EPC and/or disclosure within the meaning of Article 83 EPC is to be found, however, for only a very small proportion of the compounds and methods claimed. In the present case, the claims so lack support, and the application so lacks disclosure, that a meaningful search over the whole of the claimed scope is impossible. Consequently, the search has been carried out for those parts of the claims which appear to be supported and disclosed, namely those parts relating to the compounds and methods which disclose catalysts that comprise the elements molybdenum, vanadium, tellurium and niobium.

All examples given in the description exclusively relate to mixed oxides of Mo, V, Te and Nb. Present claims 1 to 11, however, include mixed oxides of a large number of other elements having an entirely different chemical nature. The description does not give any reason whatsoever as to why an extrapolation to these elements should be possible. It is generally accepted in the field of catalysis that there is no predictability of the catalytical performance of a particular combination of elements. Therefore, the above mentioned lack of support arises.

PARTIAL EUROPEAN SEARCH REPORT

Application Number
EP 02 25 2866

DOCUMENTS CONSIDERED TO BE RELEVANT			CLASSIFICATION OF THE APPLICATION (Int.Cl.7)
Category	Citation of document with indication, where appropriate, of relevant passages	Relevant to claim	
D,A	EP 0 608 838 A (MITSUBISHI CHEM IND) 3 August 1994 (1994-08-03) * claims 1,9 *	9-11	
			TECHNICAL FIELDS SEARCHED (Int.Cl.7)

**ANNEX TO THE EUROPEAN SEARCH REPORT
ON EUROPEAN PATENT APPLICATION NO.**

EP 02 25 2866

This annex lists the patent family members relating to the patent documents cited in the above-mentioned European search report.
The members are as contained in the European Patent Office EDP file on
The European Patent Office is in no way liable for these particulars which are merely given for the purpose of information.

11-02-2003

Patent document cited in search report	Publication date	Patent family member(s)	Publication date
US 6063728 A	16-05-2000	JP 11047598 A	23-02-1999
		DE 19835247 A1	11-02-1999
		EP 0895809 A1	10-02-1999
		US 6036880 A	14-03-2000
		US 6143916 A	07-11-2000
		JP 11239725 A	07-09-1999
EP 1254707 A	06-11-2002	CN 1382523 A	04-12-2002
		EP 1254707 A1	06-11-2002
		US 2002183199 A1	05-12-2002
EP 1192985 A	03-04-2002	BR 0104320 A	04-06-2002
		CN 1346698 A	01-05-2002
		EP 1192985 A1	03-04-2002
		JP 2002159856 A	04-06-2002
		US 2002038052 A1	28-03-2002
EP 0962253 A	08-12-1999	BR 9901568 A	21-12-1999
		CN 1236672 A	01-12-1999
		EP 1260495 A2	27-11-2002
		EP 0962253 A2	08-12-1999
		JP 2000024501 A	25-01-2000
		US 6180825 B1	30-01-2001
		US 2001049336 A1	06-12-2001
EP 0529853 A	03-03-1993	BR 9203080 A	30-03-1993
		CN 1069723 A ,B	10-03-1993
		DE 69208565 D1	04-04-1996
		DE 69208565 T2	26-09-1996
		EP 0529853 A2	03-03-1993
		JP 3168716 B2	21-05-2001
		JP 5279313 A	26-10-1993
		SG 42912 A1	17-10-1997
		US 5472925 A	05-12-1995
		US 5281745 A	25-01-1994
		JP 3331629 B2	07-10-2002
		JP 5208136 A	20-08-1993
EP 0608838 A	03-08-1994	JP 3237314 B2	10-12-2001
		JP 7010801 A	13-01-1995
		JP 3334296 B2	15-10-2002
		JP 6279351 A	04-10-1994
		DE 69402567 D1	22-05-1997
		DE 69402567 T2	27-11-1997
		EP 0608838 A2	03-08-1994
		US 5380933 A	10-01-1995

EPC FORM P0459

For more details about this annex : see Official Journal of the European Patent Office, No. 12/82